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Properties of inorganic compounds pdf

Physical properties of inorganic compounds. Give 2 properties of inorganic compounds. Thermodynamic properties of inorganic compounds. Properties of inorganic compounds pdf. Optical properties of inorganic compounds. Nomenclature and main properties of inorganic compounds. Chemical properties of inorganic compounds. Four properties of inorganic compounds.

It has just imaginable that the materials based on the physical properties of the solid inorganic compounds have carried out these decisive roles in current technology and industry. You might think that this field belongs more to the physics of materials. However, apart from the theories of physical properties, the contribution of chemistry and chemists to the preparation of materials and their structural analysis was greater than that of other branches of science. The science of the material is the application of the fundamental physical properties of the materials such that the basic theories and their applications converge. Therefore, detecting applications in these fields, the contours of research themes and their purposes can be understood. The important inorganic materials are interviewed from the chemical point of view by focusing on the relationship between preparation and insulation and physical structure and physical properties. A semiconductor is an electric conductor with electrical resistance in the range of about 10⁴ to 10⁸ ohms. A typical semiconductor is a super-high silicon that is manufactured on a large scale and is widely used for information processing devices such as computer and energy conversion devices such as solar cells. VLSI (large-scale integrated circuits) are printed on wafer made with single crystals of almost defective silicon with diameters not less than 20 cm, prepared by polycrystalline silicon from the Czochralski method. The memory chips with a very high level of integration and highly efficient computer chips have recently been realized. In a short periodic table, silicon is an IV group element and has four valence electrons. Although silicon semiconductors currently represented 90% or more of all semiconductors, isoelectronic compounds 1: 1 of compounds II-VI or III-V groups form compound semiconductors and are also used for optical or ultra-ad electronic devices high speed. For example, ZNS, CDS, GaS, INP, etc. The typical compound semiconductors and the development of technologies to grow the individual crystals of these materials are remarkable. The light emission (LED) or semiconductor lasers are important compound semiconductor applications. Since the subtle films of the compound semiconductors are made by MBE (Beam Beam Epitaxy) or MOCVD (deposition of the metalloorganic chemical steam), special organometallic compounds, such as Trimethyl Gallium GA (CH₃)₃ and trimethylarsenico as (CH₃)₃, which has almost Found a small application are now industrially used. Exercise (PageDex {3}) Which compounds other than the examples provided are suitable as semiconductors? Answer ZNO, CDS, PBS, GAP and superconductivity superconductivity is a phenomenon of zero electrical resistance below a critical temperature, TC, and was discovered in 1911 by Kamerlingh Onnes (1913 Nobel Prize for Physics), which has succeeded To liquefy in helium during his experiments to measure the electric resistance of mercury at ultra low temperatures. About 1/4 of the elements, as NB (TC = 9.25 K), in, SN and PB behave as superconductors and over 1000 alloys and intermetallic compounds are also superconductors, but only the NB-TI League (TC = 9.5 K) and NB3SN (TC = 18 K) finds the application. NB3SN, NB3GE, V3GA, etc., type A-15 compounds are cubic compounds, in which transition metal atoms are aligned in chains, and the interatomic distances are more short than those of crystalline loose metal, raising density of the states of the conduction band and the critical temperature, TC, of the compound. Figure ("PageDex {4}):- structure of YBA2CU3O7-X. Among the inorganic compound superconductors, chalcogenide mxmo6x8 compounds (x = s, if, te, em = pb, sn, etc.) of molybdenum called chevrel phases and super-temperature super-temperature derivatives à €

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